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Section 10.3 – Percent Composition and Chemical Formulas. The percent by mass (percent composition) of an element in a compound is the number of grams of the element divided by the mass in grams of the compound multiplied by 100%. % mass of element =  $\frac{\text{mass of element}}{\text{mass of compound}} \times 100$ .

## **Chapter 10 - Chemical Quantities**

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## **CHAPTER 2/CHAPTER 10/CHAPTER 11**

Chapter 10 "Chemical Quantities" Vocab.  
the SI unit representing  $6.02 \times 10^{23}$   
representative particles of a substance.  
the temperature and pressure at which  
one mole of gas occupies a volume of  
22.4 L. equal volumes of gases at the  
same temperature and pressure contain  
equal numbers of particles.

## **Quia - Chapter 10 "Chemical Quantities" Vocab**

In this lesson, we'll explore what an  
enzyme is, what an enzymatic reaction  
is and what happens to the rate of a  
reaction if you add more enzymes  
Chapter 10 chemical quantities

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10 1 kg 103 g 368 Chapter 10 Chemical  
Calculations and Chemical Equations.  
10.1 Equation Stoichiometry 369 The  
ratio of moles of P 4O 10 to moles of P  
(which came from the subscripts in the  
chemical formula, P 4O 10) provided the  
key conversion factor that allowed us to  
convert

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## **Chapter 10 Chemical Calculations and equations**

Chemical Quantities THE MOLE AND  
QUANTIFYING MATTER 10.1 The Mole: A  
Measurement of Matter Essential  
Understanding The mole represents a  
large number of very small particles.  
Reading Strategy Frayer Model The  
Frayer Model is a vocabulary  
development tool. The center of the

## **Chemical Quantities - AP Biology**

Chapter 10 Chemical Quantities 91.  
SECTION 10.1 THE MOLE: A  
MEASUREMENT OF MATTER (pages  
287-296) This section defines the mole  
and explains how the mole is used to  
measure matter. It also teaches you how  
to calculate the mass of a mole of any  
substance. Measuring Matter (pages  
287-289)

## **SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)**

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short answer and problem solving questions on

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•However, they can all be written by chemical equations that chemists use to describe chemical reactions. •In every chemical reaction, atoms are rearranged to give new substances. -Just like following a recipe, certain ingredients are combined (and often heated) to form something new. Chapter Seven 7.1 -Equations for Chemical Reactions

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